

1 *Supporting Information*

2 ***Title: The impacts of pollution sources and temperature on the light absorption***
3 ***of HULIS are revealed by UHPLC-HRMS/MS at the molecular structure level***

4 *Tao Qiu^a, Yanting Qiu^b, Yongyi Yuan^a, Rui Su^c, Xiangxinyue Meng^b, Jialiang Ma^d, Xiaofan Wang^b,*
5 *Yu Gu^a, Zhijun Wu^{b,*}, Yang Ning^a, Xiuyi Hua^a, Dapeng Liang^{a,*}, Deming Dong^a*

6 ^a Key Lab of Groundwater Resources and Environment of the Ministry of Education, College of
7 New Energy and Environment, Jilin University, Changchun 130012, China

8 ^b State Joint Key Laboratory of Environmental Simulation and Pollution Control, College of
9 Environmental Sciences and Engineering, Peking University, Beijing 100871, China

10 ^c State Key Laboratory of Inorganic Synthesis and Preparative Chemistry, College of Chemistry,
11 Jilin University, Changchun 130012, China

12 ^d Department of Chemistry, Aarhus University, 8000 Aarhus C, Denmark

13 *Corresponding Authors:

14 Dapeng Liang (liangdp@jlu.edu.cn) and Zhijun Wu (zhijunwu@pku.edu.cn).

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Text S1. Method optimization, validation, and application of LC-MS

25 representative standards (see **Table S1**) were selected to optimize the LC-MS method. These standards commonly found in HULIS samples from biomass burning (Hettiyadura et al., 2021; Huo et al., 2021; Li et al., 2020a; Wang et al., 2017a, b, 2019), coal combustion (Huang et al., 2023; Huo et al., 2021; Ma et al., 2023), secondary formation (Jaoui et al., 2023; Kojima et al., 2010; Lee and Lane, 2010; Li et al., 2021), and others (Yan et al., 2023; Zhou et al., 2022). Additionally, the elemental composition (i.e. CHO, CHN, CHON, CHOS, and CHONS) and structural differences (i.e. aromatic, aliphatic, and alicyclic structures) of HULIS were also taken into account for such selection.

Selection of mobile phase additives. We compared the detection efficiency of 25 standard compounds using four common mobile phase additives, i.e. formic acid (FA), acetic acid (AcOH), ammonium formate (AmFm), and ammonium acetate (AmAc) (Huang et al., 2020; Huo et al., 2021; Jemal et al., n.d.; Jin Yang et al., 2012; Liang et al., 2013; Song et al., 2022). **Figure S8A** showed the peak area of targeted analytes at the optimal concentration of four mobile additives, and indicated that AmFm were conducive to the signal enhancement of a few alkaloids (i.e. 1tBI and 3PAAE) and FA improved the most in the detection efficiency of some N-containing compounds (i.e. 1234THQ, 2N135T, and 4M5NC), while the more abundant CHO, CHN, CHON, CHOS, and CHONS species were detected at the highest efficiency by AcOH (detailed in **Table S2**). As a result, AcOH was selected as the best mobile phase additives in this study. The addition of AcOH increased the protonation yield of analytes but too high concentration of electrolytes will result in the loss of analyte signals (Kamel et al., 1999; Pan et al., 2006). In **Figure S8C**, it can be seen that an increase in the percentage of acetic acid caused an obvious increase of analyte signals, but acetic acid concentration higher than 0.05% v/v may cause a major decrease of analyte signals (e.g. ES, 910PheQ). It was worth mentioning that we also compared the difference between water / methanol and water / acetonitrile as mobile phases and the detection efficiency of all targeted analytes in water / acetonitrile was much lower than in water / methanol (**Figure S8B**). Therefore, we recommended that the mobile phase should be consistent with the sample solvent.

Elution rate of mobile phase. For C18 and other reversed-phase chromatographic columns, the separation performance is much sensitive to the elution rate (Lesellier et al., 2011; Louisi et al., 2006). **Figure S8D** indicated that the lower elution rate was beneficial to improving the sensitivity of analytes and thus facilitate the detection of trace HULIS species in atmosphere. However, the elution rate too low will cause the number of the theoretical plates and the column efficiency to plunge, resulting in the agglomeration of peaks (**Figure S10**). Therefore, we selected 0.2 mL min⁻¹ as the optimal elution rate.

Optimization of mass spectrometric parameters. We also compared the influence of mass spectrometric parameters on the detection efficiency of targeted analytes. Compared with chromatography, optimization of mass spectrometric parameters impacted less on the analyte signal when the analyte has been fully ionized. It will not be described in detail here and can be summarized (**Figure S9**) that (1) the higher the voltage was, the analyte signals increased but slightly. (2) An excessively wide mass range caused the signal intensity and stability of some compounds to decrease significantly, such as the signal of 4NP, 2N135T, and 3NSA with the condition of m/z 50-750 were much less than other conditions. (3) Elevating the ion transfer tube and vaporizer temperature within appropriate limits could significantly promote the ionization process of analytes. (4) The effect of different ratios of sheath gas and auxiliary gas on the ionization of analytes was neither regular nor significant.

Method validation. We calculated the method detection limits (MDLs,) of the optimized LC-MS method for 25 standard compounds based on 3 times of signal-to-noise ratio. The MDLs of these standards ranged from 0.001-0.48 ng mL⁻¹ (converted into 0.0008-0.38 ng m⁻³ in atmospheric concentration, **Table S2**). These MDLs were lower than MDLs from previous methods (Chow et al., 2015; Nyiri et al., 2016; Tao et al., 2024), which facilitated the discovery of non-targeted HULIS species with low concentrations.

All selected standards detected using this method showed good linear relationship between concentration and peak area (R^2 ranged from 0.9928-0.9997, **Table S2**). The matrix effects (described in **Text S2**) of most standards were below 15% (**Table S2**), indicating that ion enhancement or suppressions had little impact on the results. In addition, the precision (relative standard deviation, RSD) was less than 10% for all standards.

Compounds identification and quantification. The obtained data analysis was performed with the Compound Discoverer 3.3 software to generate reasonable molecular formulas and match fine structures to MS/MS data. The numbers of atoms restriction of formula were 1-40 for C, 1-100 for H, 0-40 for O, 0-6 for N, and 0-2 for S. The formulas were also constrained by setting $0.3 \leq H/C \leq 3.0$, $0 \leq O/C \leq 1.2$, $0 \leq N/C \leq 1.0$, and $0 \leq S/C \leq 0.8$. (Kind and Fiehn, 2007) All of the mathematically formulas for each peak were performed with a mass tolerance of ± 5 ppm and peak areas more than three times of the blank sample. Three curated spectral databases, mzcloud library database, ChemSpider library database, and CFM-ID (<https://cfmid.wishartlab.com>) were applied to screen suspect candidates of structure. According to the Schymanski's confidence level (CL), these candidates were divided into confirmed structures (CL1), probable structures (CL2), and tentative candidates (CL3) (Schymanski et al., 2014). We showed two examples to illustrate the derivation processes of candidates in **Figure S11**.

Aiming to address the contents of identified compounds without authentic standards, a semi-quantitative strategy was used as following description: quantify the identified compounds based on the external standard solutions of surrogate standards that has similar structure with each compound (Nguyen et al., 2014; Nozière et al., 2015). There was a simple semi-quantitative example that applied the standard curve of 4-methyl-5-nitrocatechol to quantify not only 4-methyl-5-nitrocatechol, but also 3-methyl-5-nitrocatechol and 3,4-dimethyl-5-nitrocatechol. This strategy was based on the experiment experience that substances with similar structures and same concentration will generate close signal intensities of mass spectra (e.g. Pyridine-3-carbaldehyde and Imidazole-2-carboxaldehyde shown in **Figure S8D**).

The application of this method in subsequent field samples successfully identified 264 species. CHO, CHON, and CHOS compounds were dominated species, which accounted for 47.9% - 66.8%, 18.0% - 42.1%, and 5.7% - 18.8% of all identified compounds, respectively. In terms of CHO and CHON categories, aromatic compounds were the main fraction (detailed in **Table S6**), including phenols, oxygenated polycyclic aromatic hydrocarbons, nitrophenols and other atmospheric organic species. These compounds were strongly related to fuel combustion. All of CHOS compounds were characterized by ion fragment m/z 96.9595 in the MS/MS spectra and were therefore identified as organosulfates (OSs). The majority of OSs, which can be summarized as the molecular formulas of $C_nH_{2n+2}O_{4-6}S$ and $C_nH_{2n}O_{4-6}S$ ($10 \leq n \leq 18$), were most probably derived from the secondary reaction of saturated or slightly unsaturated aliphatic hydrocarbons with SO₂.

Compared with previous studies (**Table S3**) (Chen et al., 2016; Claeys et al., 2012; Fan et al., 2018; Huang et al., 2020; Ma et al., 2020; Mukherjee et al., 2020; Stone et al., 2009; Voliotis et al., 2017; Zhang et al., 2022),

we not only identified specific structure of HULIS species but also significantly improved the mass contribution of identified species. These included the newly identified 226 species, constituting 40.2% to 51.5% of the total HULIS mass. Among them were herbicides (e.g. diisopropylamine), Nicotine degradation products (e.g. 3-pyridylacetic acid), dyes (e.g. 2-methoxy-5-methylaniline and Michler's ketone), industrial chemicals (e.g. pyridine-2,6-dimethanol), personal care products (e.g. lauroylsarcosine), and explosive residue (e.g. 3,5-dinitrophenol), highlighting the method's capability in detecting unknown HULIS species. Overall, this optimized method can effectively detect the major species of atmospheric HULIS and provide invaluable insights into their composition and diversity.

Text S2. Matrix effect evaluation.

It has been widely accepted that matrix constituents can influence analyte ionization, calibration curves, and signal responses, which was referred to the matrix effect (ME) (Antignac et al., 2005; Nasiri et al., 2021; Pizzutti et al., 2016). The ME (%) can be calculated via the following equation:

$$ME (\%) = \left(\frac{\text{peak area std in blank matrix}}{\text{peak area std in solvent}} - 1 \right) \times 100 \quad (S1)$$

In this equation, ME (%) refers to the percentage of matrix calculated by the peak area data from the analytical solutions in blank matrix extract and in organic solvent (methanol). If ion suppression occurs, it will cause negative values while ion enhancement will result in positive values.

Text S3. Calculation method of glass transition temperature.

The glass transition temperature (T_g) of HULIS was an important parameter for evaluating its phase state. According to a series of developed parameterizations (Li et al., 2020b), this study calculated the T_g of HULIS as follows:

$$T_{g,CHO} = (n_C^0 + \ln(n_C))b_C + \ln(n_H)b_H + \ln(n_C)\ln(n_H)b_{CH} + \ln(n_O)b_O + \ln(n_C)\ln(n_O)b_{CO} \quad (S2)$$

$$T_{g,org} = \sum_i w_i T_{g,i} \quad (S3)$$

$$T_g = \frac{(1-w_{org})T_{g,w} + \frac{1}{k_{GT}}w_{org}T_{g,org}}{(1-w_{org}) + \frac{1}{k_{GT}}w_{org}} \quad (S4)$$

The glass transition temperature of a certain compound under dry condition could be calculated based on its molecular formula. For example, Equation (S2) illustrated the calculation method of the glass transition temperature of CHO compounds, the value of n_C , n_H , and n_O represented the number of carbon, hydrogen, and oxygen atom, the coefficients n_C^0 , b_C , b_H , b_{CH} , b_O , and b_{CO} were 12.13, 10.95, -41.81, 21.61, 118.96, and -24.38. For CHON and CHOS compounds, they could be also calculated by parameterizations developed by (Li et al., 2020b). The bulk glass transition temperature of organic matters was equal to the weighted average of the glass transition temperature of each compound ($T_{g,i}$, as equation (S3)). In equation (S4), T_g represented the glass transition temperature of organic-water mixtures, k_{GT} was the Gordon-Taylor constant (suggested to be 2.5) for organic-water mixtures, w_{org} was the mass fraction of organic matters in the model of organic-water mixtures and the water content could be estimated using the effective hygroscopicity parameter. (Petters and Kreidenweis,

2007)

Text S4. Strong BrC chromophores screening process based on partial least squares regression (PLS) model

The PLS model in Simca 14.1.0 software was used to screen strong BrC chromophores in this study. The principle of screening was based on the following assumption: the mass absorption efficiency of HULIS is equal to the weighted average of the mass absorption efficiency of its each component, which can be summarized as the formula (S5).

$$MAE_{365,HULIS} = \sum_i w_i MAE_{365,i} \quad (S5)$$

Where $MAE_{365,HULIS}$ represents the MAE_{365} of HULIS, $MAE_{365,i}$ represents the MAE_{365} of the number i compound in HULIS samples, w_i represents the mass contribution of the number i compound to the HULIS samples.

This model has been successfully applied to screen the key BrC molecules in organic aerosols by (Kuang et al., 2023). In this study, we screened the strong BrC chromophores of HULIS as the following principles: (1) variable importance for the projection (VIP) > 1; (2) regression coefficient (coeff) > 0; (3) belonging to aromatic CHON compounds; (4) detection ratio > 90%. The screening process of the PLS model was detailed as follows.

Firstly, we imported the data and set the mass fraction of each compound as the X-variable and the MAE_{365} of each sample as the Y-variable. Then the X-variables were scaled.

	Y	2	3	4	5
Seco	MAE365	2-Aminophenol	2-Amino-4-cresol	2-Aminophenol	2-Amino-6-ethenylphenol
2	1.7545	0.00202597	0.00236814	8.45221e-005	0.000671589
3	1.97561	0.00154217	0.00211147	4.81844e-005	0.00047892
4	2.00382	0.00208761	0.00282579	0.000170308	0.000785397
5	1.92956	0.00162652	0.00229347	0.000121301	0.000673864
6	1.94559	0.00134286	0.00192335	7.71324e-005	0.00042192
7	2.02636	0.00143632	0.00194451	6.43054e-005	0.000519571
8	1.67712	0.00160901	0.00210841	4.65389e-005	0.000915948
9	1.6909	0.00183289	0.00238642	9.92225e-005	0.000674004
10	1.66049	0.00115296	0.00111547	2.73244e-005	0.00057547
11	1.72392	0.00136792	0.00175027	3.98911e-005	0.000580758
12	1.33487	0.000931725	0.00142349	1.30103e-005	0.000475339
13	1.76278	0.0016694	0.0020968	4.12437e-005	0.000701264
14	2.02735	0.00187798	0.00219455	7.6645e-005	0.000671627
15	2.11604	0.0016111	0.00168208	1.65286e-005	0.000592785
16	1.90169	0.00157396	0.00227983	4.79459e-005	0.000742677
17	1.62277	0.00128983	0.0016381	4.00108e-006	0.000741656
18	1.79419	0.00176751	0.00204895	3.81712e-005	0.000647409
19	1.76634	0.00193312	0.00199995	9.91086e-005	0.000690729

Figure S1. Schematic diagram of dataset and data normalization process.

Secondly, autofit the model using cross validation rules to determine the number of the significant components, and add component until R^2Y and Q^2 decrease. It is recommended that R^2Y and Q^2 are more than 0.5 and the difference between them is less than 0.3.

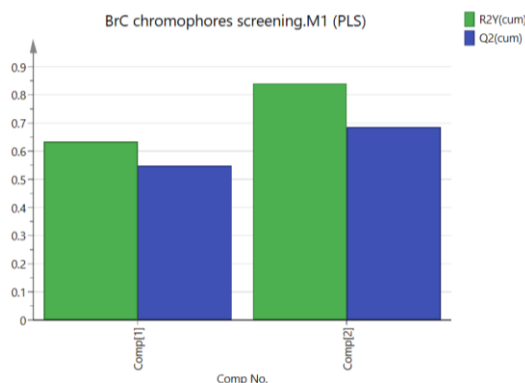


Figure S2. Components, R^2Y and Q^2 of the automatically fitted PLS model.

Thirdly, perform response permutation testing of the selected Y-variable. It is recommended that the intercept of R^2 and Q^2 should be less than 0.3 and 0.05 respectively.

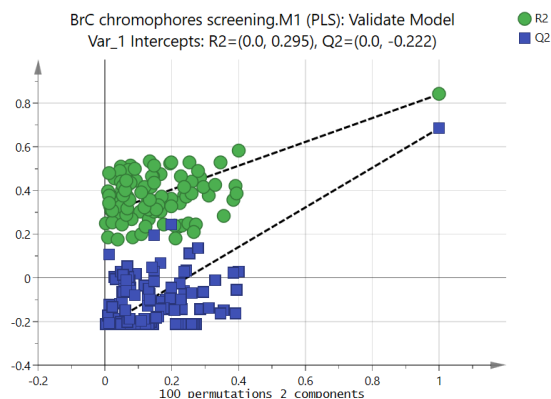


Figure S3. Response permutation test results of PLS model.

Finally, export the variable importance for the projection (VIP) and coefficient value of each compound calculated by the PLS model. The VIP value explains the importance of the independent X-variable to the dependent Y-variable. The coefficient greater than zero indicates that X-variable is positively correlated with Y-variable.

1	2	3	4	1	2	3	4
Var ID (Primary)	Var ID (Var. Sec. ID:1)	M1.VIP[2]	2.44693 * M1.VIP[2]cvSE	Var ID (Primary)	Var ID (Var. Sec. ID:1)	M1.CoeffCS[2](Var. 1)	2.44693 * M1.CoeffCS[2](Var. 1)cvSE
2 Var 69	4-Methyl-2-nitrobenzene-1,3-diol	1.74443	0.649467	2 Var 2	2-Aminophenol	0.00258402	0.0515663
3 Var 83	4,6-Dimethyl-2-nitrobenzene-1,3-diol	1.72873	0.884661	3 Var 3	2-Amino-4-cresol	-0.00136921	0.0448629
4 Var 96	4-Ethyl-3-(hydroxymethyl)-2-nitrophenol	1.66093	1.07548	4 Var 4	2-Aminophenol	0.00542233	0.0280378
5 Var 64	3,6-Dimethyl-2-nitrophenol	1.5491	0.661609	5 Var 5	2-Amino-6-ethenylphenol	-0.0147225	0.0647897
6 Var 76	1-(2-Methyl-5-nitrophenyl)ethanol	1.54396	0.50327	6 Var 6	3-Pyridylacetic acid	-0.0244832	0.0188326
7 Var 73	1-(4-Methyl-3-nitrophenyl)ethan-1-one	1.51693	0.435311	7 Var 7	2-Methoxy-5-methylaniline	-0.00625155	0.0411295
8 Var 82	2,3-Dimethyl-5-nitrobenzene-1,4-diol	1.49341	0.504874	8 Var 8	Pyridine-2,6-dimethanol	-0.0121636	0.0253226
9 Var 87	1-(2-Ethenyl-4-nitrophenyl)ethan-1-ol	1.47666	0.566255	9 Var 9	2-Amino-4-cresol	-0.0441551	0.0505053
10 Var 91	3,6-Diethyl-4-nitrophenol	1.47376	0.393309	10 Var 10	2-Formylindole	0.013755	0.0325866
11 Var 61	3-Nitrocatechol	1.45683	0.643635	11 Var 11	8-Hydroxyquinoline	-0.0115118	0.0429324
12 Var 102	2-Methyl-4-(4-nitrophenyl)but-3-en-1-ol	1.45337	0.82825	12 Var 12	1-Methyl-1H-indol-2-ol	-0.00275549	0.0229108
13 Var 67	3-Methyl-5-nitrocatechol	1.43318	0.380435	13 Var 13	Methyl-3-aminobenzoate	0.00570588	0.0461954
14 Var 68	4-Methyl-5-nitrocatechol	1.41632	0.415623	14 Var 14	3-Hydroxy-4-ethyl-1,2-benz	0.00266976	0.0382523
15 Var 80	2-(1-Hydroxyethyl)-4-nitrophenol	1.40602	0.238148	15 Var 15	1-Methoxy-2-(pyridin-3-yl)h	-0.0349914	0.0255668
16 Var 9	2-Amino-4-cresol	1.38269	1.77447	16 Var 16	1H-Indene-3-carboxamide	0.0216069	0.0178253
17 Var 15	1-Methoxy-2-(pyridin-3-yl)ethan-1-ol	1.36681	0.871277	17 Var 17	1-(1H-Indol-3-yl)ethan-1-ol	-0.0103808	0.0259645
18 Var 93	2-(1-Hydroxyethyl)-3-methyl-4-nitrophenol	1.36656	0.305176	18 Var 18	2-Methyl-1,2-dihydrophthal	0.00931875	0.033515
19 Var 81	3,4-Dimethyl-5-nitrocatechol	1.31515	0.308317	19 Var 19	6-Amino-3,4-dihydro-1H-q	0.00147662	0.0324627
20 Var 6	3-Pyridylacetic acid	1.28813	0.456705	20 Var 20	7-Hydroxy-1,2-dihydroquin	0.0095173	0.0377358
21 Var 95	3-(1-Hydroxyethyl)-2-methyl-4-nitrophenol	1.28539	0.457925	21 Var 21	Methyl 2-[(hydrazinecarbonyl	0.00168096	0.0547629
22 Var 50	1-(2,4-Dihydroxyquinolin-3-yl)ethan-1-ol	1.28452	0.744024	22 Var 22	3-Pyridineacetic acid ethyl	-0.0126746	0.0250778
23 Var 65	2,6-Dimethyl-4-nitrophenol	1.27123	0.963451	23 Var 23	6-Ethenyl-5-ethyl-1,2,3,4-te	0.00656912	0.0364175
24 Var 103	2-Ethyl-4,6-dimethyl-5-nitrobenzene-1,3-diol	1.26507	0.474343	24 Var 24	2-sec-Butyl-3-methoxyazobenz	-0.00910125	0.0382659

Figure S4. The VIP and coefficient list of each compound.

Text S5. Calculation method for the contribution of strong BrC chromophores to the total light absorption

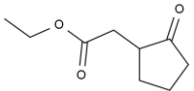
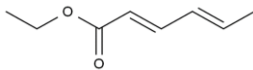
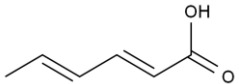
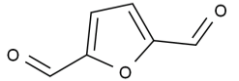
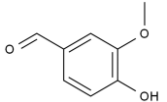
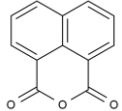
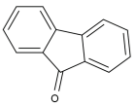
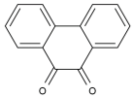
To quantify the light absorbance contribution of the identified strong BrC chromophores at a wavelength of 365 nm, we established the following calculation scheme.

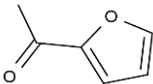
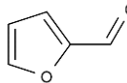
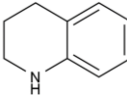
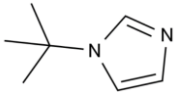
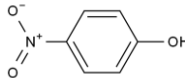
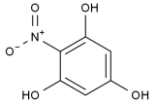
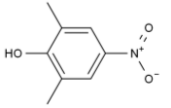
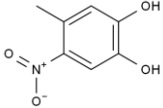
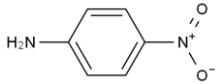
$$MAE_{strong,365} = \frac{MAE_{obs,365} \times C - MAE_{others,365} \times (C - C_{strong})}{C_{strong}} \quad (3)$$

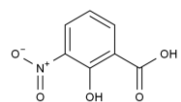
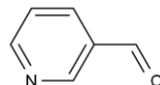
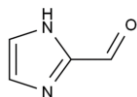
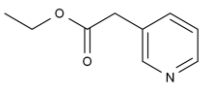
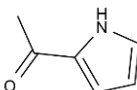
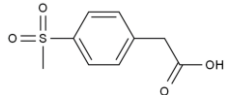
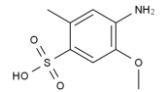
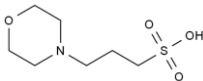
$$Light\ absorption\ contribution_{strong,365} = \frac{MAE_{strong,365} \times C_{strong}}{MAE_{obs,365} \times C} \quad (4)$$

Where $MAE_{strong,365}$ represents the MAE_{365} of strong BrC chromophores, $MAE_{obs,365}$ is the MAE_{365} of the HULIS samples calculated by formula (2), C_{strong} represents the carbon mass concentration of the strong BrC chromophores. $MAE_{others,365}$ is the MAE_{365} of components other than the strong BrC chromophores in HULIS, expressed as the intercept (set as $1.3615 \text{ m}^2 \text{ gC}^{-1}$, see **Figure 3** in **Section 3.3**) of the linear equation between the mass fraction of strong BrC chromophores with the $MAE_{obs,365}$.

Table S1. The relevant information of reference standards

Compound	Abbreviation	Structure	Formula	<i>m/z</i> & preferred ion	Category	Subcategory	Reference Sample
Ethyl-2-oxocyclopentylacetate	E2OCA		C ₉ H ₁₄ O ₃	171.1016 [M+H] ⁺	CHO	Alicyclics	
Ethyl sorbate	ES		C ₈ H ₁₂ O ₂	141.0910 [M+H] ⁺		Aliphatics	
Sorbic acid	SA		C ₆ H ₈ O ₂	113.0597 [M+H] ⁺			Biomass burning(Hettiyadura et al., 2021; Huo et al., 2021)
Furan-2,5-dicarbaldehyde	F25DC		C ₆ H ₄ O ₃	125.0233 [M+H] ⁺			
Vanillin	VAN		C ₈ H ₈ O ₃	153.0546 [M+H] ⁺			Coal combustion (Huo et al., 2021), Snowpack (Zhou et al., 2022) & Biomass burning (Huo et al., 2021; Lin et al., 2012)
1,8-Naphthalic anhydride	18NaPA		C ₁₂ H ₆ O ₃	199.0390 [M+H] ⁺		Aromatics	Secondary formation (Kojima et al., 2010), Coal combustion (Huo et al., 2021) & Biomass burning (Huo et al., 2021)
9-Fluorenone	9FLU		C ₁₃ H ₈ O	181.0648 [M+H] ⁺			Coal combustion (Huo et al., 2021)
9,10-Phenanthrenequinone	910PheQ		C ₁₄ H ₈ O ₂	209.0597 [M+H] ⁺			Secondary formation (Lee and Lane, 2010)

2-Acetylfuran	2AF		C ₆ H ₆ O ₂	111.0441 [M+H] ⁺			
2-Furaldehyde	2FA		C ₅ H ₄ O ₂	97.0284 [M+H] ⁺			
1,2,3,4-Tetrahydroquinoline	1234THQ		C ₉ H ₁₁ N	134.0964 [M+H] ⁺	CHN	Aromatics	Biomass burning (Wang et al., 2019)
1-t-Butylimidazole	1tBI		C ₇ H ₁₂ N ₂	125.1073 [M+H] ⁺			Biomass burning (Wang et al., 2017b)
4-Nitrophenol	4NP		C ₆ H ₅ NO ₃	140.0342 [M+H] ⁺	CHON	Aromatics	
2-Nitrobenzene-1,3,5-triol	2N135T		C ₆ H ₅ NO ₅	170.0084 [M-H] ⁻			Secondary formation (Jaoui et al., 2023)
2,6-Dimethyl-4-nitrophenol	26D4NP		C ₈ H ₉ NO ₃	166.0499 [M-H] ⁻			Coal combustion (Huo et al., 2021) & Biomass burning (Huo et al., 2021)
4-Methyl-5-nitrocatechol	4M5NC		C ₇ H ₇ NO ₄	168.0291 [M-H] ⁻			Secondary formation (Li et al., 2020a, 2021), Biomass burning (Li et al., 2020a) & Diesel (Yan et al., 2023)
4-Nitroaniline	4NA		C ₆ H ₆ N ₂ O ₂	139.0502 [M+H] ⁺			Ambient aerosol (Kuang et al., 2023)

3-Nitrosalicylic acid	3NSA		C ₇ H ₅ NO ₅	182.0084 [M-H] ⁻				Biomass burning (Wang et al., 2017a)
Pyridine-3-carbaldehyde	P3C		C ₆ H ₅ NO	108.0444 [M+H] ⁺				
Imidazole-2-carboxaldehyde	I2C		C ₄ H ₄ N ₂ O	97.0396 [M+H] ⁺				Secondary formation (Kampf et al., 2012)
3-Pyridineacetic acid ethyl	3PAAE		C ₉ H ₁₁ NO ₂	166.0863 [M+H] ⁺				
2-Acetyl pyrrole	2AP		C ₆ H ₇ NO	110.0600 [M+H] ⁺				
4-(Methylsulfonyl) phenylacetic acid	4MSPAA		C ₉ H ₁₀ O ₄ S	215.0373 [M+H] ⁺	CHOS	Aromatics	Snowpack (Zhou et al., 2022)	
4-Amino-5-methoxy-2-methylbenzenesulfonic acid	4A5M2MBSA		C ₈ H ₁₁ NO ₄ S	218.0482 [M+H] ⁺	CHONS	Aromatics		
3-Morpholin-4-ylpropane-1-sulfonic acid	3M4P1SA		C ₇ H ₁₅ NO ₄ S	210.0795 [M+H] ⁺		Alicyclics		

205 **Table S2.** The peak area, matrix effect of targeted analytes at the optimal concentration of four mobile additives and method detection limits (MDL), R² for standard
 206 curves, relative standard deviation (RSD) of the optimized method.

Compounds	Control		1mM AmFm		1Mm AmAc		0.05% v/v FA		0.05% v/v AcOH		MDL ng m ⁻³	R ² for standard curves	RSD %
	Peak Area	ME%	Peak Area	ME%	Peak Area	ME%	Peak Area	ME%	Peak Area	ME%			
E2OCA	2.8E+7	2.5	4.2E+7	5.8	3.2E+7	24.8	3.7E+7	1.3	5.7E+7	8.5	0.0098	0.9997	0.6
ES	1.9E+4	-1	3.2E+5	-11.1	2.3E+5	-10.3	3.8E+5	3.2	8.1E+5	-4.1	0.38	0.9940	1.5
SA	1.4E+4	6.7	1.7E+5	-19.5	1.3E+5	-20.2	3.2E+5	5.1	2.4E+5	11.6	0.10	0.9926	3.4
F25DC	2.4E+5	1.6	8.0E+5	25.7	7.3E+5	22.4	1.1E+6	0.0	1.4E+6	3.7	0.10	0.9926	4.8
VAN	1.1E+6	6.1	1.4E+6	8.3	7.4E+5	30.2	3.0E+6	6.2	4.1E+6	6.3	0.011	0.9992	2.2
18NaPA	1.1E+6	7	3.9E+6	8.2	4.5E+6	3.1	4.8E+6	4.8	6.9E+6	11.2	0.012	0.9972	1.9
9FLU	1.8E+6	8.9	2.8E+6	8.3	4.3E+6	25.1	2.7E+6	0.0	6.2E+6	5.9	0.12	0.9989	1.3
910PheQ	3.6E+7	6.6	4.4E+7	7.1	6.2E+7	23.5	3.8E+7	1.6	8.9E+7	8.8	0.0022	0.9991	5.4
2AF	2.3E+4	8.3	6.0E+5	4.5	5.1E+5	7.9	6.7E+5	4.6	6.4E+5	1.2	0.012	0.9954	3.7
2FA	2.7E+4	5.5	1.3E+5	24.5	1.1E+5	27.8	7.4E+4	19.1	1.2E+5	12.2	0.26	0.9925	5.0
1234THQ	1.2E+5	2.5	8.7E+5	-5.8	7.0E+5	-27.8	1.4E+7	0.4	1.1E+7	0.3	0.0052	0.9943	4.7
1tBI	1.5E+7	3.4	1.8E+7	21.2	1.2E+7	4.2	1.0E+7	1.8	1.0E+7	13.1	0.0048	0.9993	5.8
4NP	6.4E+4	-11.9	8.5E+4	-2.2	1.0E+5	36.4	1.2E+5	1.3	1.7E+5	-1.0	0.035	0.9934	8.6
2N135T	3.1E+6	6.6	2.0E+6	22.6	1.1E+6	19.1	6.7E+6	10.6	3.4E+6	3.1	0.039	0.9971	6.1
26D4NP	4.7E+6	1.3	1.6E+6	-1.7	3.5E+6	51.1	3.0E+6	2.8	3.4E+6	16.6	0.0075	0.9932	7.3
4M5NC	1.9E+6	12.9	3.1E+5	33.6	1.0E+5	160.9	6.2E+6	11.9	4.0E+6	0.3	0.018	0.9956	4.2
4NA	2.9E+6	0.2	4.9E+6	5.6	4.5E+6	6.0	6.0E+6	5.7	7.9E+6	-0.1	0.017	0.9989	2.1
3NSA	3.5E+5	8.2	3.1E+5	39.9	1.2E+6	39.3	6.2E+5	19.3	1.9E+5	11.9	0.37	0.9939	7.8
P3C	1.5E+6	6.7	6.1E+6	16.6	5.4E+6	19.8	8.8E+6	-0.4	1.0E+7	5.3	0.0094	0.9991	9.4
I2C	3.2E+6	2.6	7.1E+6	5.5	3.7E+6	16.8	5.3E+6	-1.7	6.8E+6	-0.5	0.016	0.9995	4.6
3PAAE	2.3E+8	7.9	4.2E+8	9.3	2.2E+8	20.4	2.5E+8	14.4	2.9E+8	13.4	0.00079	0.9979	8.7
2AP	4.9E+5	4.6	7.7E+5	4.3	6.9E+5	20.2	1.8E+6	5.9	2.6E+6	0.3	0.016	0.9964	7.8
4MSPAA	1.4E+5	2	1.9E+6	-88.9	1.7E+5	22.7	2.9E+6	21.3	3.9E+6	14.3	0.010	0.9955	9.4
4A5M2MBSA	8.1E+5	5.5	9.8E+5	3.0	1.4E+6	12.9	7.4E+5	3.9	1.3E+6	4.0	0.040	0.9928	8.9
3M4P1SA	6.4E+6	1.3	9.0E+6	0.3	8.7E+6	15.7	2.5E+6	1.7	5.2E+6	5.6	0.031	0.9980	3.4

Table S3. Identified HULIS structures in ambient aerosols

Reference	Identification Level	Numbers	Instrument	Quantification	Identification Method	Contribution to HULIS Mass
(Claeys et al., 2012)	Specific Structure	3	LC-PDA-MS	Yes	Standards Matching	0.34 - 1.2%
(Huang et al., 2020)	Specific Structure	18	LC-PDA-MS	Yes	Standards Matching	1 - 14% (light absorption contribution)
(Ma et al., 2020)	Specific Structure	16	GC-MS	Yes	Standards Matching	1.4 – 5.4%
(Fan et al., 2018)	Specific Structure	27	TMAH-GC-MS	None	Ion Fragments Matching	Unknown
(Chen et al., 2016)	Functional Group	N.A.	FTIR	Yes	Peak Wavelength Matching	100.8 - 101.1%
(Zhang et al., 2022)	Functional Group	N.A.	FTIR	None	Peak Wavelength Matching	Unknown
(Mukherjee et al., 2020)	Functional Group	N.A.	FTIR	None	Peak Wavelength Matching	Unknown
(Voliotis et al., 2017)	Functional Group	N.A.	FTIR	None	Peak Wavelength Matching	Unknown
(Stone et al., 2009)	Functional Group	N.A.	LC-MS/MS	None	Ion Fragments Matching	Unknown
This study	Specific Structure	264	UHPLC-HRMS/MS	Yes	Standards Matching & Ion Fragments Matching	38.2 – 78.1%

Table S4. Detailed chromatographic and mass spectrometric parameters used for the analysis of HULIS samples.

Instrument	Ultimate 3000 Liquid Chromatography-Orbitrap Fusion™ Tribrid™ Mass Spectrometer
LC settings	Column
	Accucore C18 100 × 2.1 mm 2.6 μm, Thermo Fisher Scientific
	Injection Volume
	3 μL
	Mobile Phases
	Phase A: ultrapure water +0.05% acetic acid Phase B: methanol +0.05% acetic acid
	Solvent Flow
	0.2 mL/min
	Column temperature
	35°C
MS settings	Gradient
	Time(min)
	%A
	%B
	0
	80
	2
	80
	17
	0
MS/MS settings	19
	0
	100
	20
	80
	20
	23
	80
	20
	Ion Source Type
MS settings	H-ESI
	Spray Voltage
	Positive: 3500 V; Negative: 3250 V
	Gas Flow (arb)
	Sheath: 35; Auxiliary: 10, Sweep: 0
	Temperature
	Ion transfer tube: 320°C, Vaporizer: 125°C
	Mass range
	60-600
	Orbitrap Resolution
MS/MS settings	120 000
	RF lens (%)
	60
	Filter criteria
	Dynamic exclusion: exclude ions after 3 times if they occur within 25 s and exclusion duration set as 100 s
	Data dependent mode
	Cycle time
	Time between Master Scans
	1 s
	Activation type
MS/MS settings	HCD
	Collision Energy (%)
	40 ± 10
	Detector type
	Orbitrap
	Resolution
	15 000

Table S5. Meteorological parameters and concentrations of pollutants in two haze events.

Parameters	Event I	Event II
PM _{2.5} ($\mu\text{g}/\text{m}^3$)	159.5 \pm 53.8	83.7 \pm 36.4
HULIS-C ($\mu\text{gC}/\text{m}^3$)	6.68	4.65
OC ($\mu\text{gC}/\text{m}^3$)	27.8	16.6
EC ($\mu\text{gC}/\text{m}^3$)	3.2	4.6
OC/EC	8.7	3.6
Na ⁺ ($\mu\text{g}/\text{m}^3$)	0.26	0.42
K ⁺ ($\mu\text{g}/\text{m}^3$)	1.26	7.31
NH ₄ ⁺ ($\mu\text{g}/\text{m}^3$)	13.88	4.49
Cl ⁻ ($\mu\text{g}/\text{m}^3$)	3.69	7.42
NO ₃ ⁻ ($\mu\text{g}/\text{m}^3$)	27.66	6.18
SO ₄ ²⁻ ($\mu\text{g}/\text{m}^3$)	11.35	7.61
SO ₂ ($\mu\text{g}/\text{m}^3$)	20.9 \pm 5.3	25.1 \pm 15.1
NO ₂ ($\mu\text{g}/\text{m}^3$)	69.5 \pm 12.6	31.5 \pm 13.3
O ₃ ($\mu\text{g}/\text{m}^3$)	14.2 \pm 10.4	33.1 \pm 19.2
CO (mg/m ³)	1.56 \pm 0.34	0.79 \pm 0.20
Relative Humidity (%)	83.1 \pm 4.6	61.9 \pm 14.0
Temperature (°C)	-10.0 \pm 4.1	-18.3 \pm 4.3
Wind Speed (m/s)	1.6 \pm 0.6	1.7 \pm 0.9

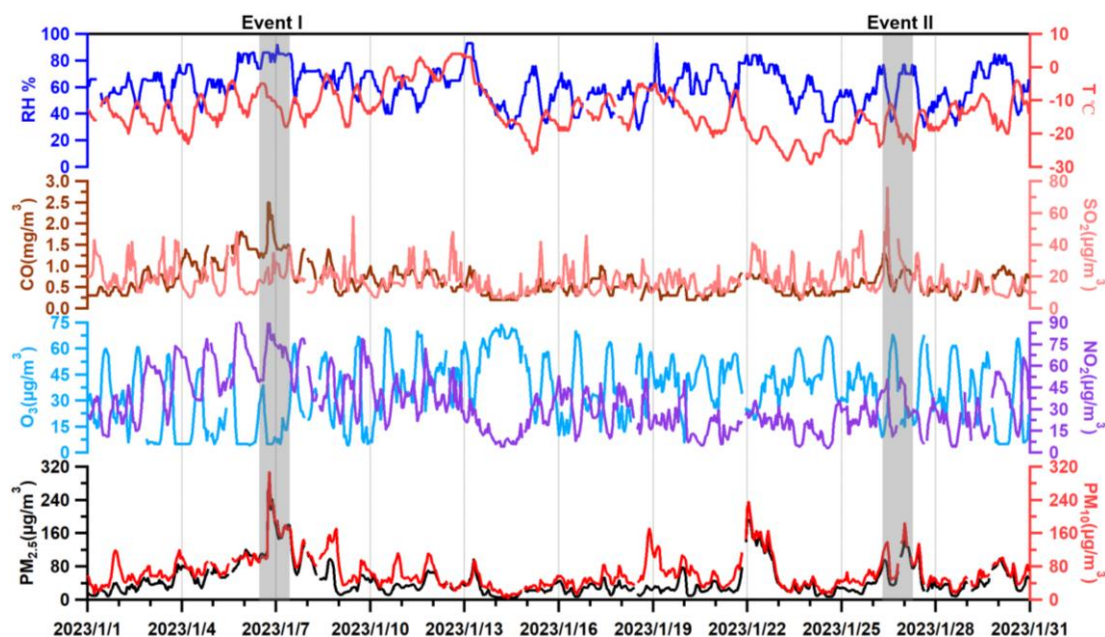


Figure S5. Time series of meteorological factors and mass concentrations of gaseous pollutants and particulate matters during sampling period.

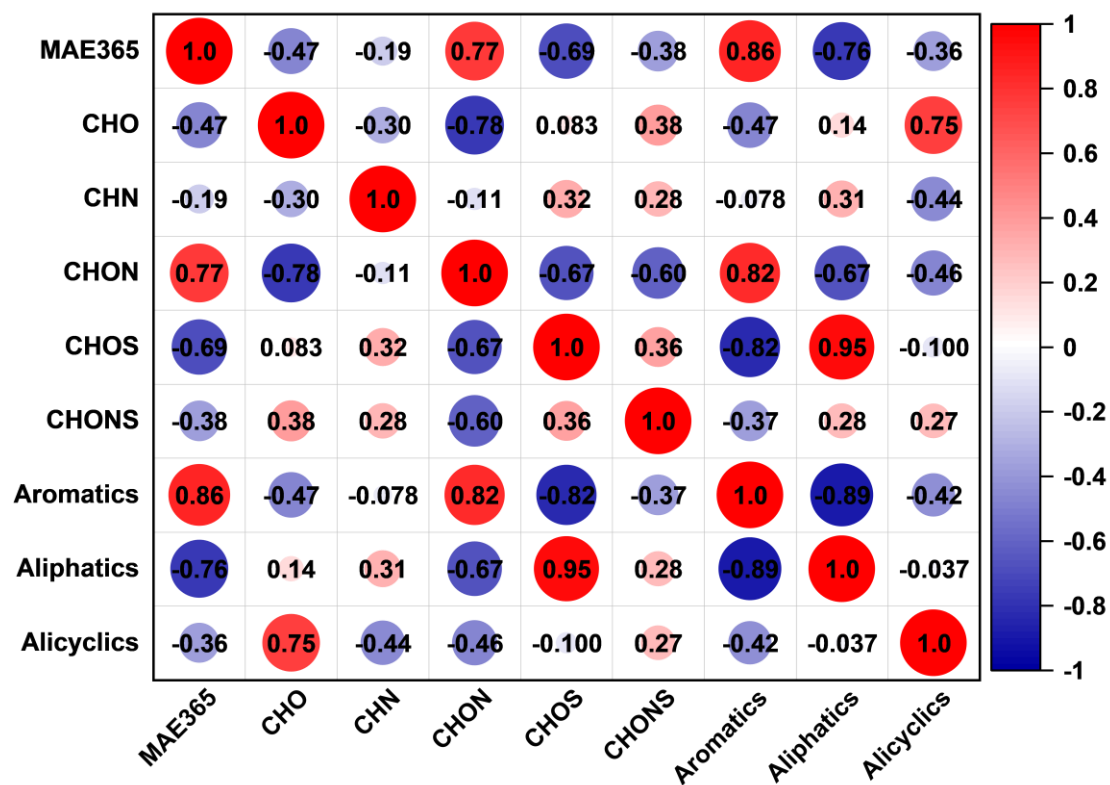


Figure S6. The Pearson's correlation matrix between elemental composition, chemical structure and MAE₃₆₅.

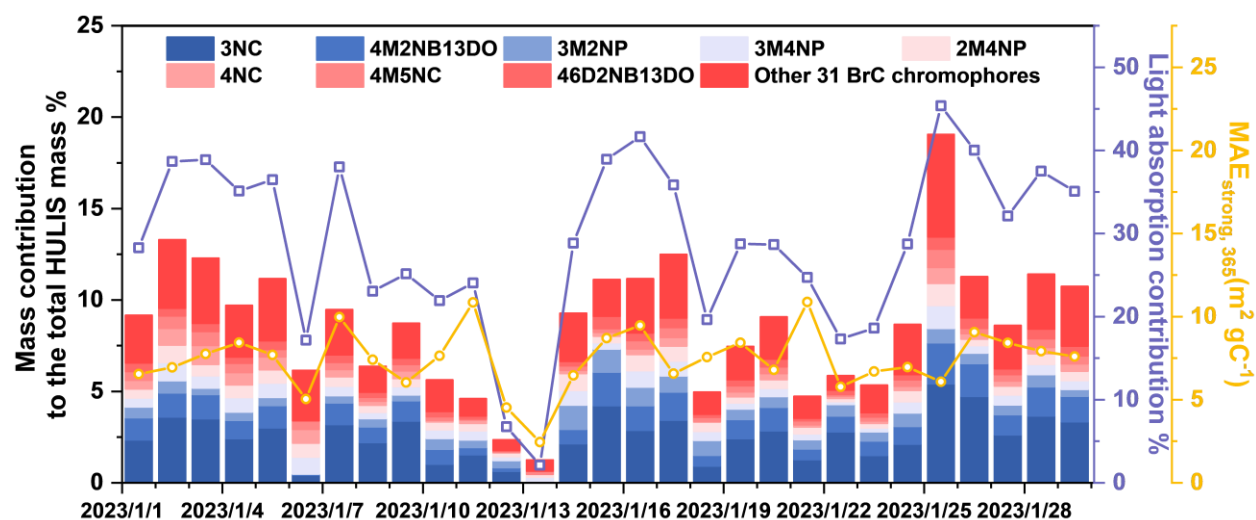


Figure S7. Temporal variations in the mass fraction, light absorption contribution, and MAE_{365} values of 39 strong BrC chromophores. The abbreviations of 3NC, 4M2NB13DO, 3M2NP, 3M4NP, 2M4NP, 4NC, 4M5NC, 26D4NP, and 46D2NB13DO represented 3-nitrocatechol, 4-methyl-2-nitrobenzene-1,3-diol, 3-methyl-2-nitrophenol, 3-methyl-4-nitrophenol, 2-methyl-4-nitrophenol, 4-nitrocatechol, 4-methyl-5-nitrocatechol, and 4,6-dimethyl-2-nitrobenzene-1,3-diol, respectively.

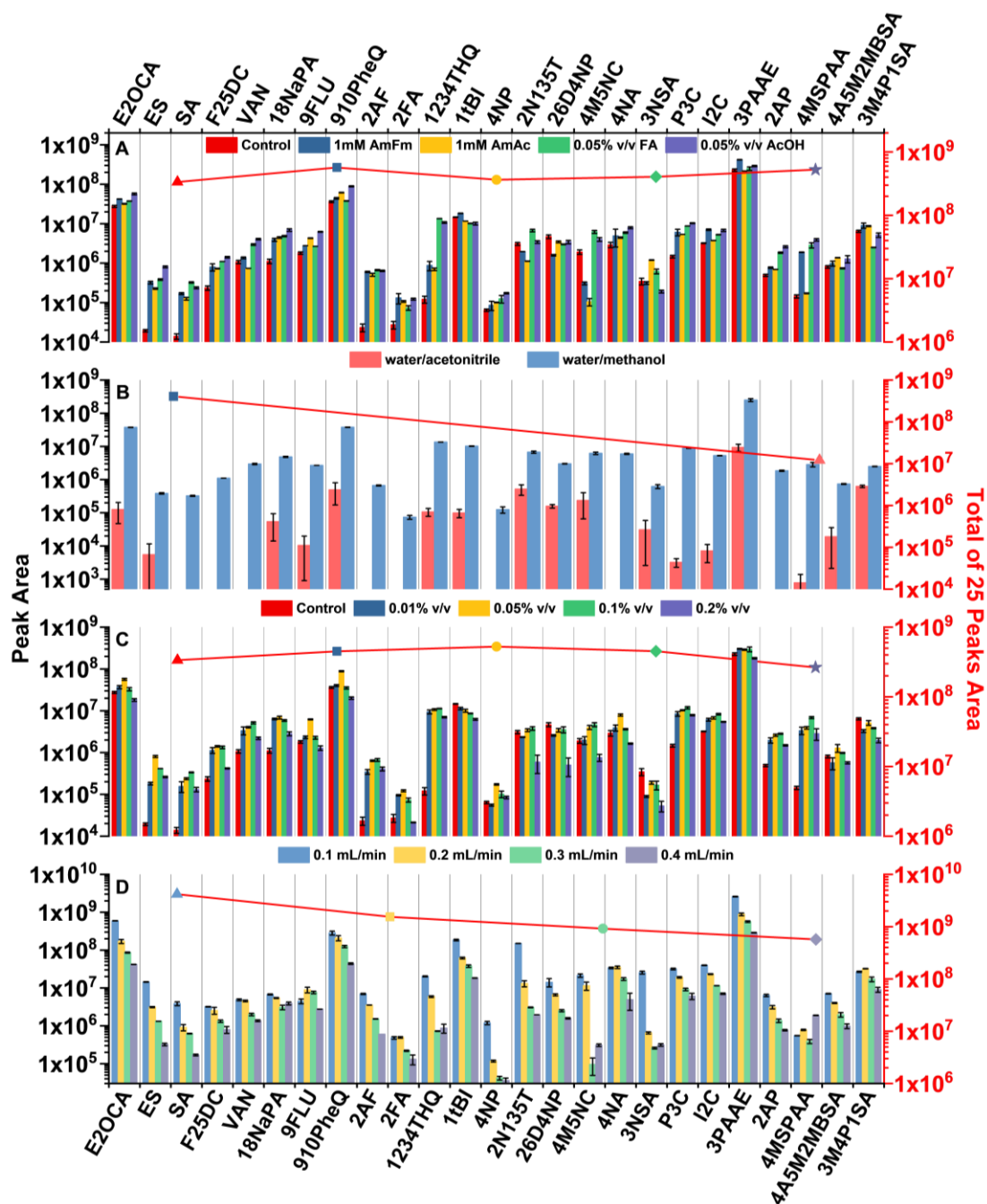


Figure S8. The peak area comparison of 25 reference standards under different type of additives (A), mobile phase (B), additive concentration (C), and flow rate (D).

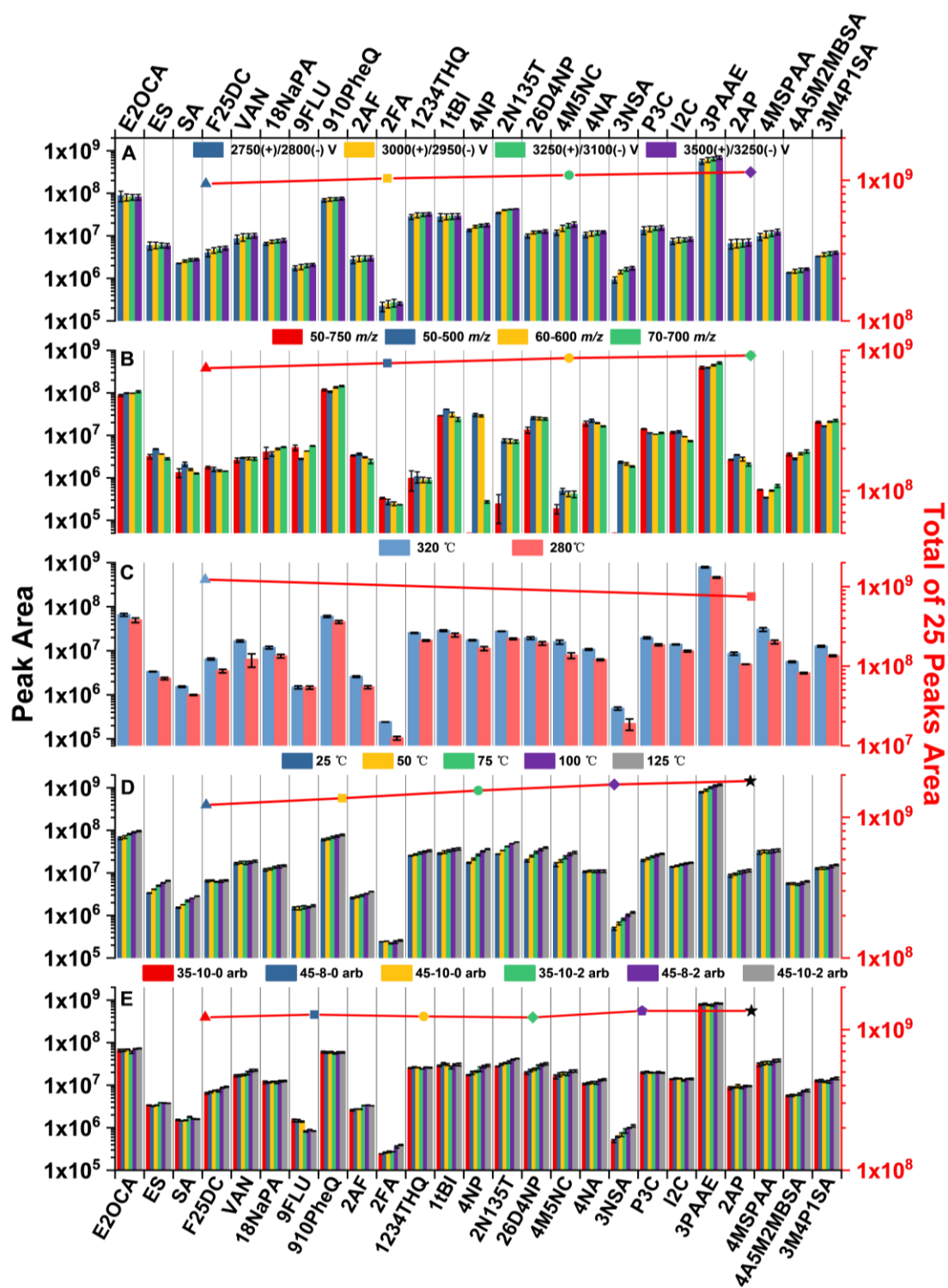


Figure S9. The peak area comparison of 25 reference standards under different voltage (A), mass range (B), temperature of ion transfer tube (C) and vaporizer (D), and the proportion of sheath-auxiliary-sweep gas (E).

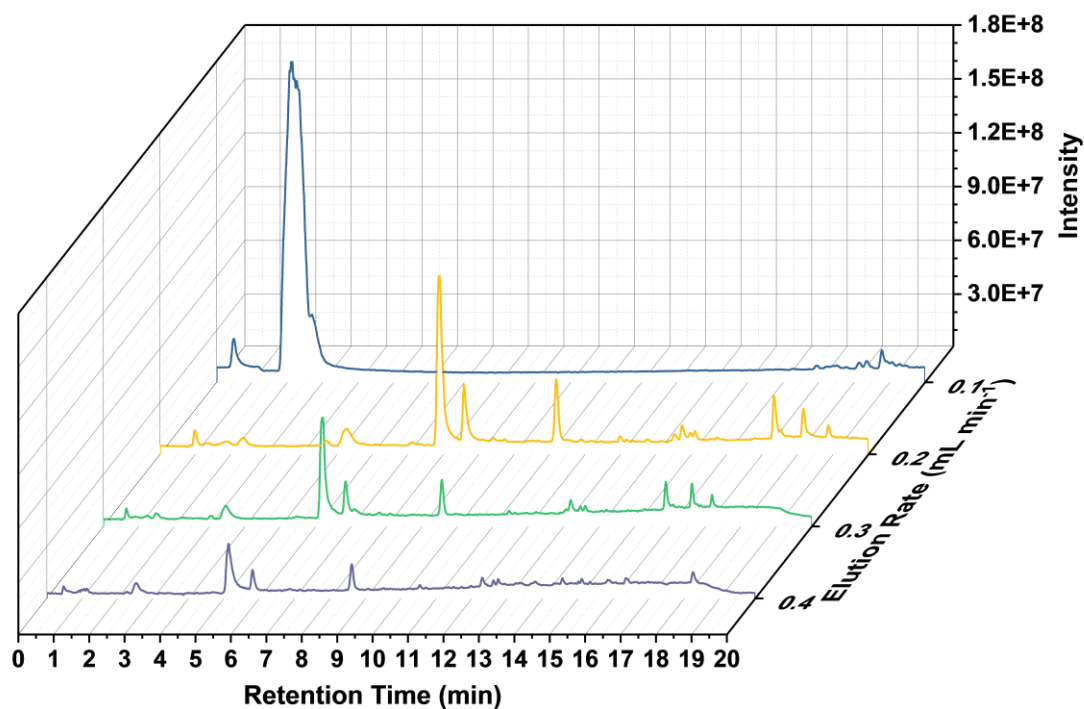


Figure S10. 3D waterfall diagram of total ions chromatogram for mixed standards at different flow rates of mobile phase.

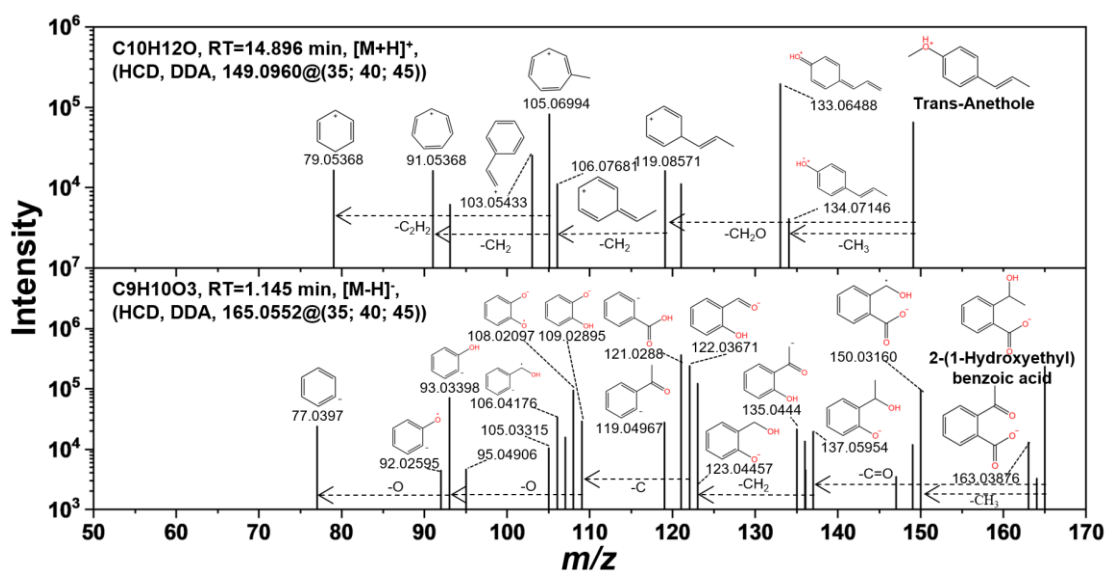


Figure S11. Schematic diagram of the MS/MS ion fragments for non-targeted HULIS species and their derivation processes.

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